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(54) Process for producing a resin having carbonate bonding.

(57) A process for producing a resin having carbonate bondings, which comprises polymerizing a hydroxydiaryl compound and phosgene in a methylene chloride organic solvent, wherein the concentration of carbon tetrachloride, if present, in the phosgene and methylene chloride satisfies the following equation:

$$A + 5B \leq 200$$

wherein A represents the carbon tetrachloride concentration (ppm) in phosgene and B represents the carbon tetrachloride concentration (ppm) in methylene chloride.

EP 0 251 586 A2

PROCESS FOR PRODUCING A RESIN HAVING CARBONATE BONDING

The present invention relates to a process for producing a resin having carbonate bondings, using methylene chloride as an organic solvent and phosgene as a starting material.

Phosgene usually contains from 250 ppm to 2000 ppm of carbon tetrachloride (CC₁₄) as disclosed in Japanese Patent Publication No. 55-14044 (1980).

5 When producing a polymer using phosgene containing CC₁₄ as a starting material and methylene chloride as a solvent, most of the CC₁₄ is dissolved in the methylene chloride and, as a result, contained in the methylene chloride solution of the polymer.

CC₁₄ in the polymer solution is distributed between the polymer and methylene chloride when they are separated in a particular ratio depending on the separation conditions.

10 Usually the separated methylene chloride is used again in industrial processes without removing the CC₁₄. Accordingly, the CC₁₄ concentration in the methylene chloride increases as the number of recycles increases. Polymers produced in this methylene chloride solvent contain acidic components, originating upon heating in pelletization and molding, which corrode dies in the molding machine when the operation thereof is stopped as disclosed in JP-A-60-81245 (1985). In this case, mold-die corrosion is prevented by 15 blending a specific ester with the molding composition.

Furthermore, the quality of the polymer is not satisfactory since the tone of pellets or molding pieces is worsened depending on the kind or amount of heat stabilizers used. It has not yet been recognized that this drawback is caused by the carbon tetrachloride present in the starting phosgene and methylene chloride.

We have studied the prevention of the formation of the acidic component upon pelletization and 20 molding of the resultant pellets, and the manner in which the tone of the pellets or molding pieces might be improved and, as a result, found that these undesirable effects are attributable to the CC₁₄ contained in the methylene chloride solution of the polymer. We have found that a resin having carbonate bondings of excellent quality, having no acidic components upon pelletization or molding and having favorable tone of 25 pellets and molding pieces can be produced by reducing the CC₁₄ concentration in the starting phosgene and methylene chloride.

Accordingly the present invention provides a process for producing a resin having carbonate bondings, which comprises polymerizing a hydroxydiaryl compound, phosgene and, optionally, a diamine compound or an acid chloride, in a methylene chloride organic solvent, wherein the concentration of carbon tetrachloride, if present, in the phosgene and methylene chloride satisfies the following equation:

$$30 \quad A + 5B \leq 200$$

wherein A represents the carbon tetrachloride concentration (ppm) in phosgene and B represents the carbon tetrachloride concentration (ppm) in methylene chloride.

Examples of a resin having carbonate bondings are thermoplastic resins such as polycarbonates, polyester carbonates and polycarbonate carbamates.

35 The polycarbonate resin includes polymers or copolymers, for example one produced from 2,2-bis(4-hydroxyphenyl)propane (bisphenol A) and phosgene.

Examples of dihydroxydiaryl compounds are bisphenol A and bis(4-hydroxyaryl)alkanes such as bis(4-hydroxyphenyl)methane,

1,1-bis(4-hydroxyphenyl)ethane,

40 2,2-bis(4-hydroxyphenyl)butane,

2,2-bis(4-hydroxyphenyl)octane,

bis(4-hydroxyphenyl)phenylmethane,

1,1-bis(4-hydroxyphenyl)phenylethane,

2,2-bis(4-hydroxy-3-methylphenyl)propane,

45 1,1-bis(4-hydroxy-3-tert-butylphenyl)propane,

2,2-bis(4-hydroxy-3-bromophenyl)propane,

2,2-bis(4-hydroxy-3,5-dibromophenyl)propane and

2,2-bis(4-hydroxy-3,5-dichlorophenyl)propane;

bis(hydroxyaryl)cycloalkanes such as

50 1,1-bis(4-hydroxyphenyl)cyclopentane and

1,1-bis(4-hydroxyphenyl)cyclohexane;

dihydroxydiaryl ethers such as

4,4'-dihydroxy diphenyl ether and

4,4'-dihydroxy-3,3'-dimethyl diphenyl ether;

dihydroxydiaryl sulfides such as

- 4,4'-dihydroxydiphenyl sulfide and
 4,4'-dihydroxy-3,3'-dimethyl diphenyl sulfide;
 dihydroxydiaryl sulfoxides such as
 4,4'dihydroxydiphenyl sulfoxide and
 5 4,4'-dihydroxy-3,3'-dimethyl diphenyl sulfoxide; and
 dihydroxy diaryl sulfones such as
 4,4'-dihydroxydiphenyl sulfone and
 4,4'-dihydroxy-3,3'-dimethyldiphenyl sulfone.

These may be used alone or as a mixture of two or more. Hydroquinone, resorcinol or 4,4'-
 10 dihydroxydiphenyl may be used in admixture with the dihydroxydiaryl compounds.

Examples of polyester polycarbonate resins are copolymers obtained by the reaction of the dihydroxydiaryl compounds mentioned above, phosgene and acid chloride (for example terephthalic chloride or isophthalic chloride) [JP-A-55-25427 (1980) and JP-A-55-38824 (1980)].

15 Examples of polycarbonate carbamates are copolymers obtained by the reaction of piperazine or dipiperidine, a dihydroxydiaryl compound and phosgene [JP-A-60-31527 (1985) and 60-35024 (1985)].

CC₁₄ is usually present in the starting phosgene at a concentration of from 250 to 2000 ppm. CC₁₄ is dissolved in the methylene chloride used in the production of a polymer and remains in the polymer solution. The polymer concentration in the polymer solution is preferably from 5 to 30% by weight, more preferably from 10 to 25% by weight.

20 The CC₁₄ in the polymer solution is distributed between the polymer and the methylene chloride upon separating them in a ratio which depends on the separation conditions. Although CC₁₄ is not usually contained in commercially available methylene chloride, since the methylene chloride separated from the polymer is recycled without separating CC₁₄ in industrial processes, the CC₁₄ concentration in methylene chloride increases as the number of recycles which it undergoes increases. Furthermore, CC₁₄ may also be produced upon separating the polymer and methylene chloride if they are heated to a high temperature.

Acidic components are formed when pelletizing a polymer produced in methylene chloride containing a great amount of CC₁₄. Upon molding such pellets, or from the resultant pellets, corrosion of the die in the molding machine is caused, and the tone of the pellets or molding pieces are not quite satisfactory.

In view of the above, we have found that the CC₁₄ concentration in the polymer solution before separating the methylene chloride and the polymer should be as low as possible.

As a method of reducing the CC₁₄ concentration in the polymer solution, the intrusion of CC₁₄ may be decreased by reducing the CC₁₄ concentration in the starting phosgene, by which the CC₁₄ concentration can be reduced to not more than 200 ppm, preferably to not more than 150 ppm.

CC₁₄ may, for example, be removed from the starting phosgene by adsorption on activated carbon or by fractional distillation utilizing the difference in boiling point between phosgene and carbon tetrachloride.

In the case of distillation, sufficient removal can be attained by a single stage distillation or distillation in a plurality of stages and this method is preferred since satisfactory results can be obtained by a relatively simple device.

The CC₁₄ content of the methylene chloride may be reduced, for example, by distillation of a portion or 40 the entire amount of methylene chloride during recycling of the methylene chloride.

To lower the CC₁₄ concentration in the methylene chloride, the dissolved oxygen concentration in the methylene chloride solution of the polymer may be decreased when evaporating the methylene chloride from the solution to leave a polymer powder.

The dissolved oxygen concentration in the methylene chloride solution of the polymer in this case is 45 preferably no more than 20 ppm, more preferably not more than 10 ppm.

The dissolved oxygen concentration in the methylene chloride solution of the polymer is usually not less than about 30 ppm when the process is carried out in air. The dissolved oxygen in the methylene chloride solution of the polymer is increased by the dissolution of oxygen from air. For instance, polymerization and washing of the polymer may be conducted in an agitation vessel. Furthermore, part of 50 the methylene chloride or the methylene chloride solution of the polymer in a storage tank may be passed to the succeeding step while the remainder is recycled to the storage tank using a centrifugal pump. Since a gas phase is present in the agitation vessel or storage tank and it is impossible to operate under completely airtight conditions, air allowed in, for example from the gland portions for the stirring, and vent pipes, is dissolved into the methylene chloride solution of the polymer due to the agitation or the recycling.

55 To decrease the dissolved oxygen in the methylene chloride solution of polymer, there are the following methods:

(1) removing dissolved oxygen from the starting materials and preventing the presence of external oxygen by conducting all of the production steps in a pressurized inert gas atmosphere;

(2) removing dissolved oxygen from the methylene chloride solution of the polymer by blowing in an inert gas before the concentration and powderization;

(3) removing the dissolved oxygen directly, for example by the addition of a reducing agent.

The method (3), using a reducing agent, is not preferred since the reducing agent, if it is a solid or a high boiling-point liquid, may remain in the polymer product and undesirably affect the quality of the product. The method (1) is not preferred since it is difficult to operate under airtight conditions. Thus method (2) is preferred. Examples of inert gases are nitrogen and argon. As an example of this method, where the nitrogen is blown at double the volume of the methylene chloride solution of the polymer into the methylene chloride solution of the polymer before the condensation and powderization, the dissolved oxygen concentration is reduced to about 1/10 and the dissolved oxygen in the methylene chloride solution of the polymer can be removed relatively simply.

While CC₁₄ may be removed by any of the methods as described above, it is necessary that the CC₁₄ concentration in the polymerization system be kept within the range defined by the following formula:

$$A + 5B \leq 200,$$

wherein A represents the CC₁₄ concentration in the phosgene (ppm) and B represents the CC₁₄ concentration in the methylene chloride (ppm); preferably A + 5B < 150 and, more preferably A + 5B ≤ 100.

It is generally satisfactory that the carbon tetrachloride concentration in the resulting polymer is not more than 20 ppm, preferably not more than 10 ppm.

The methylene chloride may be used in a mixture, for example with a chlorinated hydrocarbon such as dichlorobenzene, chloroform, tetrachloroethane, trichloroethane, dichloroethane or 1,2-dichloroethylene, or with dioxane, tetrahydrofuran, acetophenone, toluene, xylene, cyclohexane, acetone or n-heptane. Mixtures containing not less than 60% methylene chloride are preferred and 100% methylene chloride is most suitable.

To form the polymer interfacial polymerization may, for example, be used, which comprises reacting phosgene or phosgene and acid chloride with a dihydroxydiaryl compound in the presence of a diamine compound, organic solvent, water and alkali hydroxide, or a solution polymerization may be used wherein a dissolved hydroxydiaryl compound and optional diamine compound is reacted with phosgene or phosgene and an acid chloride in the presence of an organic alkali and an organic solvent as an acid acceptor.

The polymer solution after polymerization may be diluted with methylene chloride as required for washing and purification. The polymerization system referred to in the present invention may include the step of polymerizing as described above and, in the case of conducting dilution, the step of dilution as well. Impurities in the polymer solution may be removed by washing, for example with an aqueous alkali solution, an aqueous acid solution or water. Washing is generally conducted using a mixer settler or centrifugal extractor.

The polymer may, for example, be powdered by heating the polymer solution and removing methylene chloride, or by precipitation, for example by adding a non-solvent.

Since methylene chloride and non-solvents may be contained in the polymer, these solvents are removed by drying using an ordinary method.

Since the powdery polycarbonate obtained according to the present invention produces no acidic component when heated upon pelletization or molding, problems such as corrosion of the die in the molding machine can be eliminated. Thus eliminating rust deposited on the molding machine is not necessary upon restarting the molding machine after stopping. While previously a great amount of polymer has been necessary for removing rust in the molding machine, such removal is no longer required and consequently the loss of the polymer is decreased.

The tone of the resultant pellets or molding pieces is stable and favorable due to the reduction in the CC₁₄ concentration. Products of substantially satisfactory quality can thus be obtained.

The present invention is explained further in the following Examples.

50 Examples 1 - 3 and Comparative Example 1 - 3

To remove CC₁₄ contained in phosgene (concentration: 500 ppm), 2 liters of liquefied phosgene were charged into a 3 liter flask and subjected to rectification for 4 hours using a Widmer rectification pipe at a column temperature of 8°C and a reflux ratio of 1:1.

As a result, the CC₁₄ concentration in the phosgene was reduced to 5 ppm (1.66 kg) at the top, but was 1200 ppm (1.17 kg) at the bottom of the column.

A polycarbonate was prepared by using the thus resultant phosgene from the column top in accordance with an ordinary method [refer to JP-A-61-14227 (1986)]. That is, 15.3 liters of purified water, 1.03 kg (25.7 mol) of sodium hydroxide, 3.0 g of sodium hydrosulfite and 2,790 g (12.2 ml) of bisphenol A were successively added under stirring and completely dissolved in a 25 liter glass-lined vessel equipped with 5 baffle plates and having a stirrer, a phosgene blowing tube, a thermometer and a drain cock disposed at the bottom portion.

Then, 3.1 liters (7.65 mol) of an aqueous 10 wt/vol % sodium hydroxide solution and 7.65 liters of methylene chloride (CC₁₄ concentration: 5 ppm) were added. 1.38 kg (13.9 mol) of phosgene was blown in under stirring for 60 minutes while keeping the liquid temperature at 20°C by external cooling, and, 10 thereafter, 66.7 g of para-tert-butylphenol and 3.1 g of triethylamine were added. The mixture was stirred for one hour and polymerization took place.

After the completion of the polymerization and one washing of the organic phase in the lower layer with purified water, 3.0 liters of 10% phosphoric acid was added for neutralization, the mixture was washed with purified water three times and heated to distil off methylene chloride and to obtain a powdery polycarbonate. 15

The powdery polycarbonate was dried by an ordinary method and kneaded into pellets with no addition of additives such as heat stabilizer in a 20 mmØ extruder (manufactured by Tanabe Plastic Co.) (kneading temperature: 280°C, number of screw rotations: 40 rpm, kneading rate: 1.6 kg/Hr, polymer staying time: 2 min).

20 The resultant pellet was in the shape of 1.3 mmØ and 2.5 mmL (length).

The pellets were dried by an ordinary method and YI (yellow index) was measured by a color difference meter (Model TC-55D, manufactured by Tokyo Denshoku Co.). The pellet had satisfactory YI of 6.

Measurement in a metal corrosion test was then conducted by the method described below.

25 After molding 200 sheets of flat plates of 7 cm length, 4 cm width and 3.2 mm thickness by injection molding (molding temperature 300°C) using carbon steel molding dies (S55C), the molding dies after use were left at room temperature for 24 hours and the surface state of them was evaluated by visual observation. The ratings for the evaluation were as below:

Evaluation 0 : no rust

Evaluation 1 : point-corrosion

30 Evaluation 2 : partial brown rust

Evaluation 3 : entire brown rust

Evaluation result for the corrosion of the dies was satisfactory if there was no occurrence of rust, since it is necessary for practical use that the evaluation result for the corrosion is "0".

35 Measurements were also conducted for pellets obtained by adding CC₁₄ (guaranteed reagent, manufactured by Wako Junyaku Co.) to phosgene and/or methylene chloride so as to provide CC₁₄ content as shown in Table 1 below and conducting polymerization in the same manner as above. Similarly, measurements were also conducted for the case where the CC₁₄ concentration in phosgene was 200 ppm using fresh methylene chloride (CC₁₄ concentration (B) = 0). The results are as shown in Table 1 below.

For the comparison, the result in a case where the value of A + 5B exceeds 200 ppm is also shown.

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Table 1

	CCl ₄ concentration (ppm)			Die corrosion test	Pellet tone YI
	in phosgene (A)	in methylene chloride (B)	A + 5B		
Example 1	5	5	30	0	6
" 2	10	20	110	0	7
" 3	50	20	150	0	8
" 4	200	0	200	0	8
Comparative Example 1	50	50	300	1	11
" 2	50	110	600	2	15
" 3	500	20	600	2	15

Example 5

To remove CCl₄ contained in phosgene (concentration: 500 ppm), 2 liters of liquefied phosgene were charged into a 3 liter flask and subjected to rectification by using a Widmer rectification pipe at a column temperature of 8°C and a reflux ratio of 1:0.5.

As a result, a CCl₄ concentration of 50 ppm in phosgene was obtained at the column top.

A polycarbonate was produced by using the phosgene obtained using the same procedure as that described in Example 1 except that methylene chloride containing CCl₄ in a concentration of 10 ppm was used.

After the completion of the polymerization, 10.1 liters of methylene chloride prepared by adding CCl₄ to methylene chloride containing no CCl₄ so as to adjust the CCl₄ concentration to 10 ppm was added to the organic phase in the lower layer. After washing once with purified water, 3.0 liters of 10% phosphoric acid were added for neutralization and, after washing three times with purified water, methylene chloride was distilled off by heating to obtain a powdery polycarbonate.

The distilled methylene chloride was collected by cooling and 95% of the charged methylene chloride was recovered. 10% of the recovered methylene chloride was distilled off to reduce the CCl₄ concentration in methylene chloride to not more than 2 ppm. The recovered methylene chloride, the distilled methylene chloride and supplementary methylene chloride (methylene chloride corresponding to 5% of the charged methylene chloride prepared from methylene chloride containing no CCl₄ and adjusted 10 ppm by adding CCl₄) were mixed and used as the methylene chloride for polymerization and washing.

The procedures of polymerization, washing and powderization were conducted six times. The CCl₄ concentration in the recovered methylene chloride substantially reached equilibrium of 8 ppm at the fifth operation. The resulting powdery polycarbonate obtained by recycling methylene chloride six times was dried by an ordinary method and kneaded with no addition of additives such as heat stabilizers in a 20 mmØ extruder (manufactured by Tanabe Plastic Co.) into pellets (kneading temperature: 280°C, screw rotational number: 40 rpm, kneading rate: 1.6 kg/Hr, polymer staying time: 2 min).

The resulting pellets were in the shape of 1.3 mmØ × 2.5 mmL (length). The pellets were dried by an ordinary method and YI and mold die corrosion tests were carried out in the same manner as described in Example 1.

The results are shown in Table 2 below.

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Examples 6 and 7

Starting phosgene was prepared by adding carbon tetrachloride to a column top solution obtained by the same procedure as described in Example 5 (a column top solution prepared by distillation in a Widmer rectification pipe for removing CC₁₄ in phosgene) to attain a CC₁₄ concentration in phosgene as shown in Table 2 below.

A solvent was prepared by adding carbon tetrachloride to methylene chloride containing no CC₁₄ so as to adjust the CC₁₄ concentration in methylene chloride to a value as shown in Table 2 below.

Polymers were prepared using the same procedure as described in Example 5 using the phosgene and methylene chloride adjusted with the carbon tetrachloride (CC₁₄) concentration. A predetermined amount of the recovered methylene chloride was recycled after distillation and, when the CC₁₄ concentration in the recovered methylene chloride reached an equilibrium, the die corrosion by the resultant polymer and the pellet tone were evaluated.

The results are shown in Table 2 below. Satisfactory results were obtained for the die corrosion evaluation and the pellet tone in the case where the value of A + 5B is not more than 200.

Comparative Examples 4 - 6

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Polymers were prepared by the same procedure as described in Example 5 while varying the concentration of CC₁₄ in phosgene and the CC₁₄ concentration in the recycled methylene chloride. A predetermined amount of the recovered methylene chloride was recycled after distillation and when the CC₁₄ concentration in the recovered methylene chloride reached an equilibrium the die corrosion of the resultant polymer and the pellet tone were evaluated.

No satisfactory results were obtained when the value for A + 5B exceeded 200. The results are also shown in Table 2.

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Table 2

Distillation ratio for recycled methylene chloride (%)	Starting phosgene (A) (ppm)	CCl ₄ concentration		Die corrosion test	Pellet tone YI
		Recycled methylene chloride	Initial value (B) (ppm)		
Example 5	10	50	10	90	0
" 6	10	70	10	120	0
" 7	30	100	10	140	0
Comparative Example 4	10	100	15	290	1
" 5	10	200	100	71	555
" 6	5	200	100	123	815

Example 8

5 To the same glass lined vessel used in Example 1, 15.3 liters of purified water, 1.03 kg (25.7 mol) of sodium hydroxide, 3.0 g of sodium hydrosulfide and 2,790 g (12.25 mol) of bisphenol A were successively added and dissolved completely under stirring.

3.1 liters (7.65 mol) of an aqueous 10 wt/vol% of sodium hydroxide solution and 7.65 liters of methylene chloride were then added and 1.38 kg (13.9 mol) of phosgene was blown in under stirring for 60 min while keeping the liquid temperature at 20°C by external cooling. 66.7 g of para-tert-butylphenol and 3.1 g of triethylamine were subsequently added and polymerization was carried out under stirring for one hour.

10 After the completion of the polymerization, 14.1 liters of methylene chloride were added to the organic phase in the lower layer and washed once with purified water, 3.0 liters of 10% phosphoric acid were then added for neutralization and washed three times with purified water to obtain a purified methylene chloride solution.

15 When analyzing the 10% methylene chloride solution of the polymer (15 liters), the dissolved oxygen concentration was 30 ppm and the CC₁₄ concentration was 5 ppm.

When the dissolved oxygen concentration was measured after blowing a nitrogen gas at a rate of 6.0 liter/min for 10 min into the polymer solution, it was 0 ppm. The solution was charged into a nitrogen-substituted 20 liter volume autoclave made of stainless steel (SUS-316) and maintained at 180°C for 2 hours with no stirring.

20 After heat treatment, it was cooled to ambient temperature and the CC₁₄ concentration in the methylene chloride solution of the polymer was quantitatively determined to be 5 ppm by gas chromatography (GC7A manufactured by Shimazu Seisakusho).

The polymer solution of the methylene chloride was heated again and methylene chloride was distilled off to obtain a powdery polycarbonate.

25 The powdery polycarbonate was dried in an ordinary manner and kneaded into pellets with no addition of additives such as heat stabilizer in a 20 mmØ extruder (manufactured by Tanabe Plastic Co.) (kneading temperature: 280°C, number of screw rotations: 40 rpm, kneading rate: 1.6 kg/Hr, polymer staying time: 2 min).

The resulting pellets were in the form of shape of 1.3 mmØ and 2.5 mmL (length).

30 The pellets were dried by an ordinary method and YI (yellow index) was measured by a color difference meter (Model TC-55D, manufactured by Tokyo Denshoku Co.).

The pellets had a satisfactory YI value of 6.

Example 9

Dissolved oxygen was removed from the 10 wt% methylene chloride solution of a polymer prepared in the same procedure as described in Example 8 by blowing in a nitrogen gas at a rate of 6.0 liter/min for 4 minutes. The residual dissolved oxygen concentration was 5 ppm. The methylene chloride solution of the 40 polymer was charged into a 20 liter autoclave made of stainless steel under nitrogen gas and maintained at 180°C for 2 hours with no stirring in the same manner as described in Example 8.

As a result, the CC₁₄ concentration in the methylene chloride solution of the polymer after the heat treatment was 16 ppm. When the polymer solution of the methylene chloride was powdered and kneaded in an extruder, the pellet had a YI value of 7.

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Example 10

Dissolved oxygen was removed from a 10 wt% methylene chloride solution of the polymer prepared in the same procedure as described in Example 8 by blowing in a nitrogen gas at a rate of 6.0 liter/min for 2 minutes. The residual dissolved oxygen concentration was 10 ppm. The methylene chloride solution of the polymer was charged into a 20 liter autoclave made of stainless steel under nitrogen gas and maintained at 180°C for 2 hours with no stirring in the same manner as described in Example 8.

As a result, the CC₁₄ concentration in the methylene chloride solution of the polymer after the heat treatment was 25 ppm. When the polymer solution of the methylene chloride was powdered and kneaded in an extruder, the pellet had a YI value of 8.

Comparative Example 7

5 A methylene chloride solution of 10 wt% polymer concentration and 30 ppm dissolved oxygen concentration prepared in the same procedure as described in Example 8 was charged in the same 20 liter autoclave made of stainless steel (not under nitrogen) used in Example 8 and maintained at 180°C for 2 hours with no stirring.

As a result, the CC₁₄ concentration in the methylene chloride solution of the polymer after heat treatment was 65 ppm and the pellets obtained by powderizing and drying the polymer solution and then kneading in an extruder had a poor YI value of 13.

10 The results of Examples 8 to 10 and Comparative Example 7 are shown in Table 3.

Table 3

	Heat treatment temperature (°C)	Dissolved oxygen concentration (ppm)	CC ₁₄ in methylene chloride solution		Pellet tone	YI
			before heat treatment (ppm)	after heat treatment (ppm)		
Example 8	180	0	5	5		6
" 9	"	5	"	16		7
" 10	"	10	"	25		8
Comparative Example 7	"	30	"	65		13

Claims

- 40 1. A process for producing a resin having carbonate bondings, which comprises polymerizing a hydroxydiaryl compound, phosgene and, optionally, a diamine comopund or an acid chloride, in a methylene chloride organic solvent, wherein the concentration of carbon tetrachloride,if present, in the phosgene and methylene Chloride satisfies the following equation:
- $$A + 5B \leq 200$$
- 45 wherein A represents the carbon tetrachloride concentration (ppm) in phosgene and B represents the carbon tetrachloride concentration (ppm) in methylene chloride.
2. A process according to claim 1 wherein the phosgene used as a starting material contains not more than 200 ppm of carbon tetrachloride.
- 50 3. A process according to claim 1 or 2 wherein the methylene chloride is recycled.
4. A process according to claim 3 wherein the carbon tetrachloride concentration is reduced in the recycled methylene chloride by distillation in the recycling system.
5. A process according to any one of claims 1 to 4 wherein the carbon tetrachloride concentration in the phosgene used as a starting material is reduced by distillation.
- 55 6. A process according to any one of claims 1 to 5 wherein the carbon tetrachloride concentration in the methylene chloride is reduced by decreasing the dissolved oxygen concentration in the methylene chloride solution to not more than 20 ppm before it is removed by distillation to obtain the resin.

7. A process according to any one of claims 1 to 6 wherein the dissolved oxygen concentration of the methylene chloride is reduced to not more than 20 ppm by blowing in an inert gas.

8. A process according to any one of claims 1 to 7 wherein the resin having carbonate bondings is a polycarbonate.

5 9. A process according to any one of claims 1 to 8 wherein the carbon tetrachloride concentration in the resin having carbonate bondings is not more than 20 ppm.

10. A process according to claim 9 wherein the carbon tetrachloride concentration in the resin having carbonate bondings is not more than 10 ppm.

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(54) **Process for producing a resin having carbonate bonding.**

(57) A process for producing a resin having carbonate bondings, which comprises polymerizing a hydroxydiaryl compound and phosgene in a methylene chloride organic solvent, wherein the concentration of carbon tetrachloride, if present, in the phosgene and methylene chloride satisfies the following equation:

$$A + 5B \leq 200$$

wherein A represents the carbon tetrachloride concentration (ppm) in phosgene and B represents the carbon tetrachloride concentration (ppm) in methylene chloride.

EP 0 251 586 A3



EUROPEAN SEARCH REPORT

EP 87 30 5423

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl.4)
D, A	PATENT ABSTRACTS OF JAPAN, vol. 9, no. 218 (C-301)[1941], 5th September 1985; & JP-A-60 81 245 (IDEMITSU SEKIYU KAGAKU K.K.) 09-05-1985 * Abstract * ---	1	C 08 G 63/62
A	DE-A-1 943 803 (IDEMITSU KOSAN CO., LTD) * Claims 1-12 * ---	1	
A	FR-A-2 456 761 (GENERAL ELECTRIC) * Claim 1; page 5, table I * -----	1	
The present search report has been drawn up for all claims			
Place of search	Date of completion of the search	Examiner	
THE HAGUE	27-10-1988	DECOKER L.	
CATEGORY OF CITED DOCUMENTS			
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document		T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document	